
CBSE TEST PAPER-01
CLASS - IX Science (Atoms and Molecules)

1. Atomic radius is measured in nanometers and [1]
(a) $1\text{nm} = 10^{-10}\text{m}$ (b) $1\text{m} = 10^{-10}\text{nm}$
(c) $1\text{m} = 10^{-9}\text{nm}$ (d) $1\text{nm} = 10^{-9}\text{m}$
2. Symbol of Iron is :- (a) Ir (b) I (c) Fe (d) None of these [1]
3. Atomicity of chlorine and Argon is [1]
(a) Diatomic and Monoatomic (b) Monoatomic and Diatomic
(c) Monoatomic and Monoatomic (d) Diatomic and Diatomic
4. Molecular mass of water (H_2O) is (a) 18g (b) 8g (c) 33g (d) 34g [1]
5. State law of conservation of Mass? [2]
6. Define Valency? Find the Valency of oxygen and Aluminum. [2]
7. Calculate the molar mass of Na_2SO_4 and $CaCO_3$? [2]
8. Write the chemical formula for [2]
a) Calcium Phosphate b) Magnesium Hydroxide c) Aluminum chloride.
9. State the Postulates of Dalton Theory? [3]
10. Find the percentage of water of crystallization in $FeSO_4 \cdot 7H_2O$. [3]
11. 2.42g of copper gave 3.025g of a black oxide of copper, 6.49g of a black oxide, on reduction with hydrogen, gave 5.192g of copper. Show that these figures are in accordance with law of constant proportion? [3]
12. A compound was found to have the following percentage composition by mass [3]
 $Zn = 22.65\%$, $S = 11.15\%$, $H = 4.88\%$, $O = 61.32\%$. The relative molecular mass is 287g/mol. Find the molecular formula of the compound, assuming that all the hydrogen in the compound is present in water of crystallizations.