

30/12
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VRITI GUPTA
7/1/20

classmate
Date _____
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IS MATTER AROUND US PURE!

9th

Substance \rightarrow it consists of a single type of particles. They are classified as elements or compounds.

Mixtures are intermixing of two or more ^{elements} particles in any proportion to form a compound.

Solution \rightarrow it is a homogeneous mixture of particles of low solution are so small that cannot be seen with naked eyes. eg. Sugar in water.

Concentration \rightarrow the amount of solute that of a solution has been dissolved in a given amount of solvent or solution.

Suspension \rightarrow it is a heterogeneous mixture. The particles can be seen with naked eyes. Suspension is unstable. eg. Chalk powder and water.

Colloid \rightarrow it is a heterogeneous mixture. The individual particles cannot be seen with naked eyes. They are quite stable. For eg. Milk.

Element \rightarrow it is a simple substance which cannot be broken into small particles by chemical reaction.

Compound \rightarrow a substance made up of two or more elements eg. Water (H_2O).

Saturated solution \rightarrow it is a solution in which the maximum amount of solute has been dissolved at a given temperature.

Solubility \rightarrow the ability of a solute to dissolve in a solvent to form a solution.

VRITI GUPTA
IX - A



Class 9th

Is Matter Around Us Pure

Maximum Marks: 44

Q1) Define:

(1x12=12)

- a) Substance
- b) Mixture
- c) Solution
- d) Concentration of a solution
- e) Suspension
- f) Colloid
- g) Element
- h) Compound
- i) Saturated solution
- j) Solubility
- k) Unsaturated solution
- l) Tyndall Effect

Mammoth Brains
Where the best brains meet.

Q2) What we call "pure milk" is actually a mixture. Explain

(2)

Q3) Differentiate between: (Giving examples)

(2x4=8)

- a) Heterogeneous and Homogeneous Mixtures
- b) Mixtures and Compounds
- c) Compounds and Elements
- d) Sol, Solution and Suspension

Q4) Do solid and gaseous solutions also exist? If yes, give examples.

(2)

Q5) Enlist the properties of the following:

(2x3=6)

- a) Solution
- b) Suspension
- c) Colloid

Q6) Which separation technique will you apply for the separation of the following?

(1x6=6)

- a) Dye from Ink
- b) Mixture of sodium chloride and ammonium chloride
- c) Pigments from an extract of flower petals.